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12.2: Mole Ratios - Chemistry LibreTexts

Mole ratios are used as conversion factors between products and reactants in stoichiometry calculations. For example, in the reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$ The mole ratio between O₂ and H₂O is $\frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}}$. The mole ratio between H₂ and H₂O is $\frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}}$.

Example:

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I also know that balanced chemical equations are necessary for the proper use of mole ratios. Activator: I ask a student to show their work or share

their coefficients. They should be: $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. $2\text{NaBr} + \text{CaF}_2 \rightarrow 2\text{NaF} + \text{CaBr}_2$. $\text{HNO}_3 + \text{NaHCO}_3 \rightarrow \text{NaNO}_3 + \text{H}_2\text{O} + \text{CO}_2$.
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